



UDC 621.791.927.5 : 669.017.165

<https://doi.org/10.17073/1997-308X-2023-4-59-70>

Research article

Научная статья



Transformation of the nickel aluminide alloy structure through the application of a heat-resistant coating using oscillation electrode surfacing

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Abstract. This study considers the formation of an alloyed nickel aluminide structure through automatic electric arc surfacing employing an oscillating electrode composed of composite wire. The arc transversely traverses the weld pool surface at a frequency denoted as f . In comparison to conventional surfacing techniques, this process either displaces the crystallization front alongside the weld pool (at $f = 1.3$ Hz) or stabilizes it (at $f \geq 2$ Hz) throughout the cross-sectional area of the coating layer. We have conducted an investigation into the evolution of alloy structures resulting from surfacing. Notably, we have observed that the regions with concentrations of eutectic nickel-aluminum are particularly susceptible to structural alterations. The formation of particle clusters, which is contingent upon heat dissipation conditions near the crystallization front, leads to the development of layered texture regions. Our findings reveal that following 50 thermal cycles (heating to 1100 °C, cooling to 25 °C), the alloy's hardness becomes independent of subsequent thermal cycles, consistently maintaining a level 34–35 HRC. The highest resistance of the surfaced metal to thermal fatigue cracks is achieved when its structure exhibits an optimal γ -solid solution (relatively ductile) to nickel-aluminum cooling martensite ratio, corresponding to the Ni_2Al phase. The thermal conditions necessary for producing such a structure are elucidated by the gradual cooling of the crystallized metal from elevated temperatures when $f \geq 2.8$ Hz. An analysis of changes in oxidative wear, estimated by mass loss, during thermal fatigue tests conducted at a metal heating temperature of 1100 °C revealed the superiority of the studied alloy over industrial alloys based on nickel and cobalt.

Keywords: electric arc surfacing, oscillating electrode, nickel aluminide, thermal cycle, metal structure, thermal fatigue, oxidation resistance

Acknowledgements: This research received support from the Russian Science Foundation, Grant No. 23-13-00354 (<https://rscf.ru/project/23-13-00354/>).

For citation: Zorin I.V., Lysak V.I., Kharlamov V.O., Fastov S.A. Transformation of the nickel aluminide alloy structure through the application of a heat-resistant coating using oscillation electrode surfacing. *Powder Metallurgy and Functional Coatings*. 2023;17(4):59–70. <https://doi.org/10.17073/1997-308X-2023-4-59-70>

Трансформация структуры сплава на основе алюминид никеля в процессе наплавки колеблющимся электродом термостойкого покрытия

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Аннотация. Работа посвящена исследованию формирования структуры легированного алюминид никеля в процессе автоматизированной электродуговой наплавки колеблющимся электродом (плавящейся композиционной проволокой), при котором дуга совершает поперечные перемещения с частотой f по поверхности расплава сварочной ванны. Использование такого приема в сравнении с традиционной технологией наплавки позволяет перемещать вместе с расплавом сварочной ванны фронт кристаллизации (при $f = 1,3$ Гц) или стабилизировать его (при $f \geq 2$ Гц) в поперечном сечении наплавляемого металла. Изучена эволюция структуры наплавленных сплавов. Установлено, что наиболее структурно чувствительной фазой являются участки сосредоточения никель-алюминидной эвтектики, скопления частиц которой, в зависимости от условий теплоотвода вблизи фронта кристаллизации, образуют участки слоистой текстуры. Показано, что после 50 теплосмен (нагрев до 1100 °C, охлаждение до 25 °C) твердость исследуемого сплава перестает зависеть от последующего термического циклирования и сохраняется неизменной на уровне 34–35 HRC. Наибольшая стойкость наплавленного металла к появлению трещин термической усталости обеспечивается при формировании в его структуре близкого к оптимальному соотношения относительно вязкого, легированного железом и другими элементами γ -твердого раствора и никель-алюминидной мартенситной фазы, состав которого соответствует Ni_2Al -фазе. Термические условия получения такой структуры обусловлены замедленным охлаждением закристаллизовавшегося металла с высоких температур при достижении $f \geq 2,8$ Гц. Анализ изменения сопутствующего термоусталостным испытаниям окислительного изнашивания (оцениваемого потерей массы) при температуре нагрева металла 1100 °C показал преимущества исследуемого сплава над промышленными сплавами на основе никеля и кобальта.

Ключевые слова: электродуговая наплавка, колеблющийся электрод, алюминид никеля, термический цикл, структура, термическая усталость, стойкость к окислению

Благодарности: Исследование выполнено за счет гранта Российского научного фонда № 23-13-00354 (<https://rscf.ru/project/23-13-00354/>).

Для цитирования: Зорин И.В., Лысак В.И., Харламов В.О., Фастов С.А. Трансформация структуры сплава на основе алюминид никеля в процессе наплавки колеблющимся электродом термостойкого покрытия. *Известия вузов. Порошковая металлургия и функциональные покрытия*. 2023;17(4):59–70. <https://doi.org/10.17073/1997-308X-2023-4-59-70>

Introduction

Developed during the mid-20th century, weld overlay represents an efficient technique for extending the operational lifespan of products exposed to temperatures as high as 1000 °C. To enhance the resistance of the cladding layer against high-temperature wear, iron, nickel, and cobalt alloys are commonly employed [1–3]. These alloys have undergone extensive research. Conversely, nickel aluminide-based heat-resistant alloys for surfacing have garnered substantial practical interest. Among these, Ni_3Al and NiAl intermetallics are particularly promising [4; 5].

Both domestic and international nickel aluminide casting alloys exhibit improved thermal stability [4; 6], elevated high-temperature and fatigue strength [7; 8], and resistance to high-temperature oxidation [4; 9].

Despite the significant potential of these alloys, their application in cladding is constrained due to the high sensitivity of nickel aluminide to heating and cooling rates, primarily owing to the elevated γ' - Ni_3Al phase content, which diminishes plasticity and hinders stress relief during welding [10]. Consequently, it is imperative to maintain a low-temperature gradient at the crystallization front in the weld pool. Meng Zhang et al. [11] have reported achieving a similar outcome by constructing a vertical “wall” of Ni_3Al alloy via a multilayer weld overlay technique under argon shielding, incorporating multiple reheating cycles by subsequent layers, and controlling the cooling rate (1–2 °C/s).

In contrast to the “gentle” thermal deformation cycle of argon arc surfacing using a non-consumable electrode, semi- or fully automated weld over-

lay employing a consumable electrode significantly enhances both productivity and the quality of the surfacing layer. However, as highlighted by Sorokin L. et al. [12], the conditions for metal structure formation become more complex due to the higher heat input and the larger melt pool volume, resulting in increased weld stress [12].

Various methods are currently employed to control heat input during welding processes. These methods encompass the use of pulse [13] and modulated [14] weld current, high-frequency oscillations of the electrode [15], or, as the most common and cost-effective approach, adjusting the specific welding energy. The control of specific energy can influence the size and shape of the weld pool and, consequently, the properties of the surfacing metal in the clad alloy. However, this control is effective only within specific ranges of welding arc power and weld deposit rates [16], which are contingent on the particular alloying system [17; 18]. For Ni–Cr cladding, a relatively wide range of weld overlay rates can be applied to manage phase composition and structure size while ensuring a defect-free surfacing layer. Nevertheless, for high-carbon and boron-containing wear-resistant alloys, the applicable range of rates is considerably narrower.

Several strategies have been devised to minimize the conversion of weld energy into heat input at the crystallization front. Wu Dongting et al. [19] suggested excluding the workpiece from the welding circuit. Another method involves employing transverse oscillations of the weld arc across the weld pool surface. This latter approach to heat input control offers greater flexibility and was originally investigated by G.G. Chernyshev, B.N. Kushnirenko, and M.M. Shtrikman in the 1960s and 1970s. Today, it finds application in multilayer wear-resistant surfacing [20]. Pre-eutectic and eutectic wear-resistant alloys with high chromium and carbon content exhibit more favorable responses to transverse electrode oscillations with specific frequencies and amplitudes [17; 21; 22]. In certain instances, repeated thermal effects are applied to the crystallizing metal as the weld arc moves, potentially altering the carbide phase morphology and grain size [23]. It has been postulated that utilizing an oscillating electrode for cladding an alloy containing eutectic phases of nickel aluminide is a preferred approach for controlling the thermal conditions of structure formation, thereby enhancing the thermal resistance of the surfacing layer.

The primary objective of this study is to characterize the structure of the nickel aluminide alloy formed during surfacing under the influence of repeated thermal effects on the cladding layer.

Materials and Methods

We employed electrodes crafted from composite wire (CW) with a diameter of 3 mm. This wire was manufactured through a 6-pass drawing process using a carbide die. The wire sheath consists of a nickel strip, while the wire filler comprises aluminum, tungsten, molybdenum, powdered tantalum, and chromium. This composite wire design (Fig. 1) ensures a uniform distribution of filler components along the length of the wire. In the case of powder wires, achieving such uniformity typically requires a layer-by-layer arrangement of filler components possessing similar physical and mechanical properties. The inherent inhomogeneity in thermophysical properties across the cross-section is a common characteristic of composite structures, yet this is mitigated by effective electrical contact between the surfaces of the composite wire components, facilitating rapid heating. Taking into account the loss of filler elements due to oxidation within the weld pool, the theoretical composition of the composite wire corresponds to a stoichiometric nickel-to-aluminum mass ratio ~6.5. The inclusion of refractory alloying elements such as tungsten, molybdenum, and tantalum

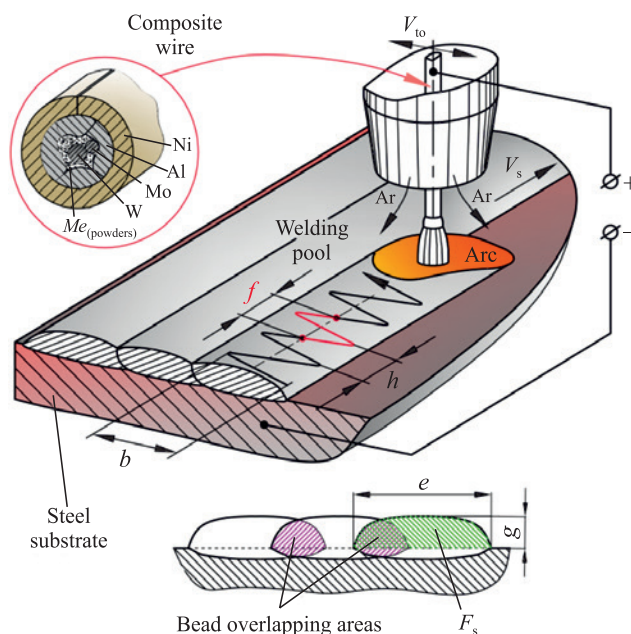


Fig. 1. Cladding process

f – frequency of the transverse arc oscillations over the weld pool;
 V_{to} – electrode transverse oscillation velocity; V_s – surfacing rate;
 h – surfacing increment; b – electrode amplitude; e – weld bead width;
 g – weld bead height; F_s – area of the surfaced part of the bead

Рис. 1. Схема формирования наплавленного покрытия

f – частота поперечных перемещений дуги по поверхности сварочной ванны; V_{to} – скорость колебаний электрода;
 V_s – скорость наплавки; h – шаг наплавки; b – размах колебаний электрода; e – ширина валика; g – высота валика;
 F_s – площадь наплавленной части валика

in the filler material strengthens the γ' -solid solid solution and enhances the stability of the γ' -phase at temperatures up to 1100 °C. Additionally, the surfacing layer contains a minimum of 4 wt. % chromium, which, when combined with aluminum, imparts robust resistance to high-temperature oxidation. As a substrate, we utilized plates measuring 12×80×170 mm, composed of the 5CrNiMn steel grade.

The resulting chemical composition (wt. %) of the nickel aluminide alloy is as follows:

Ni	Base metal	W	3.3–3.6
C	0.2–0.3	Mo	2.6–3.1
Al	8.2–8.5	Ta	2.3–2.5
Cr	4.0–4.2	Fe	12–15

Equally spaced VP 5/20 tungsten alloy thermocouples were installed within the steel substrate at depth of 1.5 mm below the surface. These thermocouples are designed to gauge the temperature at the base of the weld pool and in close proximity to the crystallization front as the base metal undergoes melting. The thermocouples were constructed using 0.6 mm diameter wire. In order to process the thermocouple signals, we utilized an LA-20USB multichannel ADC converter from Rudnev–Shilyaev, Moscow, and subsequently transferred the data to a PC. The PowerGraph™ software was employed to record the signals and generate temperature vs. time curves.

The melting point of the alloy, as determined by differential scanning calorimetry, is ~1386 °C. In order to further investigate the structure and elemental composition of the alloy, we employed an electron-ion microscope (FEI Versa 3D, USA).

In order to assess the cladding layer's resistance to thermal fatigue cracking, we conducted cyclic heating tests and monitored the number of heating-cooling cycles until the first fatigue cracks became evident. The samples, measuring 15×15×4 mm, consisted of single-layer cladding metal. Additionally, we fabricated test samples using Stoodite 6 and Hastelloy C clad alloys. These samples were then positioned within a furnace that was heated to a temperature of $t = 1100$ °C, held at this temperature for 6 min, and subsequently cooled down to $t = 25$ °C in 7–9 s. This testing protocol encompassed a total of 240 thermal cycles. In order to evaluate the oxidation resistance of the samples, we measured the mass loss of each sample after every 30 thermal cycles. These measurements were carried out using a VIBRA HT-124RCE analytical balance (Shinko Denshi Co. LTD, Japan) with a readability of 0.1 mg.

Surfacing Process

We employed an A2 Mini Master welding machine (ESAB, Sweden), which includes an electrode oscillation drive. The electrode oscillates transversely to the direction of the cladding. The oscillation drive converts the rotation of the gearmotor shaft into a reciprocating rectilinear motion of the lever connected to the electrode wire feeder. The velocity of the transverse oscillations of the electrode V_{to} spans from 16 to 50 mm/s. This velocity decreases to approximately 0.1 times V_{to} at the extremities of the electrode's oscillation amplitude (b).

The transverse oscillation velocity is synchronized with the linear surfacing velocity (V_s). Consequently, the surfacing increment (h) is automatically chosen by the welding machine to maintain the specified V_s value. An increase in V_{to} to ~50 mm/s results in a decrease in surfacing increment, causing the arc to reach the end of its transverse path within 1 s, and the oscillation frequency (f) approaches 3 Hz. Conversely, when V_{to} is reduced to 16 mm/s, the surfacing increment increases, and f shifts to its minimum value of ~1 Hz (see Fig. 1).

We introduced the weld bead shape factor (μ_b), defined as the ratio of the cross-sectional area of the bead to the area of a rectangle with dimensions equal to the bead width (e) and bead height (g) (see Fig. 1). For single-layer coatings, the bead-to-bead overlap varied in the range of 30–50 %, depending on μ_b . The steel substrate comprises 16–20 % of the surfacing layer.

In order to ensure superior weld bead quality, the surfacing velocity was set at 19 cm/min. Exceeding this value adversely affects the weld beads, resulting in inconsistencies in their shape, length, and width. As the surfacing velocity decreases (<10 m/h), the retraction of molten material into the tail of the weld pool is significantly reduced, leading to an expansion of the weld pool volume beneath the arc. This results in a shorter arc and disrupts the wire melting process. Arc stability (preventing short circuits in the arc gap with the refractory wire filler) is achieved at $U_{arc} = 27$ V. A welding current of $I_{weld} = 280 \pm 15$ A was selected to provide the arc thermal power sufficient to melt the refractory filler in the wire while preventing overheating of the drop at the wire end. Initially, the heat needed to keep the filler components in a molten state is supplied by the anode spot at the end of the refractory tungsten-molybdenum filler. When the melt entirely covers the refractory filler, the anode spot shifts to the end surface of the resulting drop. The drop separates as it descends along the surface

of the partially molten filler. Following this, a new molten filler drop begins to form around it [24].

The welding current source (DC, reversed polarity) utilized was an LAF 1001 thyristor rectifier (ESAB, Sweden). The primary surfacing process variables include:

Welding current (I), A 280 ± 15
Arc voltage (U), V 27 ± 1
Oscillation amplitude (b), mm 10
Frequency of arc transverse oscillations (f), Hz 1.3; 2.0; 2.8
Surfacing velocity (V_s), cm/min 19 ± 1
Shielding gas (argon) flow rate, l/min 25–30

Results and discussion

The analysis of the metal's heating and cooling profile during arc cladding without transverse composite wire oscillations reveals a thermal cycle with a single peak (Fig. 2, *a*). This peak corresponds to the maximum temperature reached when the melt contacts the thermocouple junction. During the cooling stage near the crystallization temperature, the cooling rate is relatively high (~ 100 °C/s). Subsequently, starting from 1300 °C, the temperature decrease gradually slows down to ~ 20 °C/s. The resulting structure of the cladding layer, formed under such non-equilibrium conditions (Fig. 3, *a*), comprises two regions with roughly equal volume fractions. One of these regions consists of the alloyed γ -solid solution, characterized by primary crystallizable dendrites. The other region contains dispersed lamellar particles of the γ' -phase. As per the established concept of the Ni_3Al alloy structure and phase state, based on the work of Kolobov Yu. et al. [4], these particles are formed through the $L \leftrightarrow \gamma + \gamma'$ eutectic reaction.

Further analysis of the eutectic concentration areas reveals significant non-uniformity in the distribution of iron and aluminum between the two primary phases. The eutectic nickel-aluminum areas exhibit the highest concentrations of aluminum and nickel (Fig. 4, *a, b*). Given that iron can form a continuous solid solution with nickel, it is primarily dissolved in the γ -phase. The presence of iron in the γ' -phase can be attributed to its ability to substitute for both nickel and aluminum in Ni_3Al .

The transverse oscillations of the electrode (at a frequency of $f = 1.3$ Hz) lead to the appearance of multiple peaks on the cooling curve (see Fig. 2, *b*). These peaks correspond to short-term temperature fluctuations caused by reheating of the cooled metal by the arc. The pattern of the peaks for 1, 2 and 3

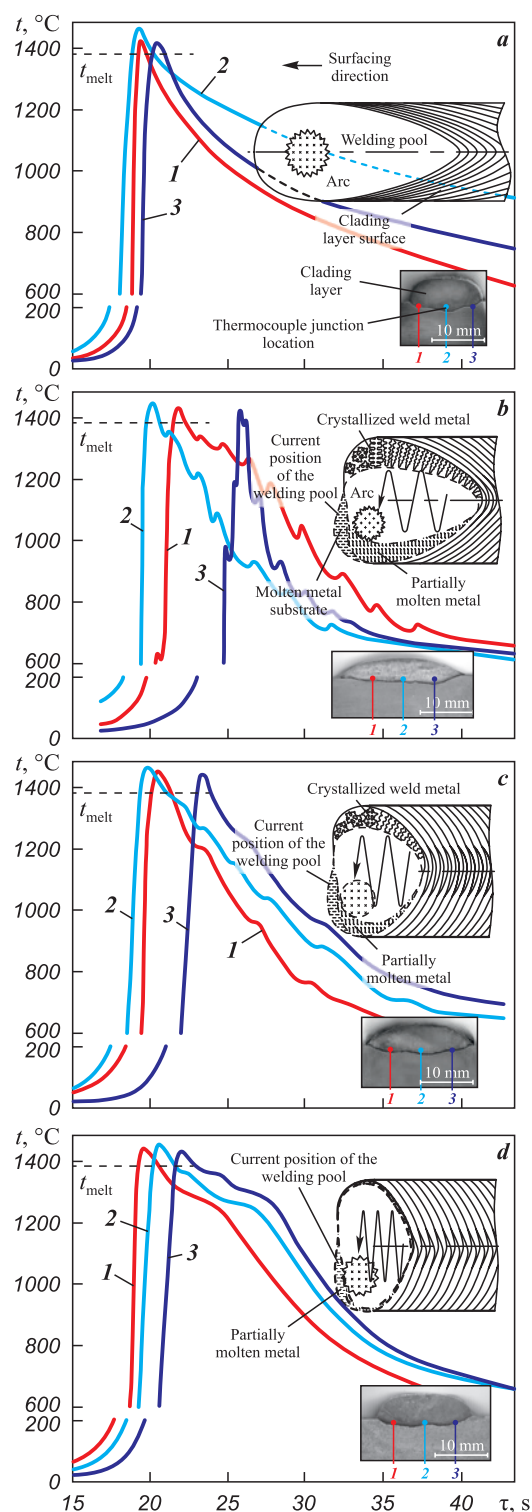


Fig. 2. Thermal cycles of the surfacing process without (*a*) and with electrode oscillations (*b–d*) at $f = 1.3$ Hz (*b*), 2 Hz (*c*), and 2.8 Hz (*d*)
1, 2, 3 are the temperature curves as recorded by the thermocouple junctions within the base metal

Рис. 2. Термические циклы процесса наплавки без колебаний электрода (*a*) и с перемещениями дуги по поверхности сварочной ванны (*b–d*) с частотой $f = 1,3$ Гц (*b*), 2 Гц (*c*), 2,8 Гц (*d*)
1, 2, 3 – кривые изменения температуры, соответствующие местам расположения спаев термопар в основном металле

thermal cycles indicate that the crystallizing metal experiences elevated temperatures for a limited duration. This phenomenon arises due to the cyclic redistribution of hot and cold areas within the weld pool as the pool moves in response to the transverse oscillations of the weld arc. However, when the weld pool tracks the heat source (the arc), it remelts the previously cooled metal, and crystallization recommences. This cyclic process results in a chaotic orientation of crystallite growth and the interweaving of crystallites, as observed in the lower and partially middle sections of the beads. The structure of the upper (working) bead (Fig. 3, *b*) displays a high degree of dispersion within the regions containing the eutectic γ' -phase. The size of such regions does not exceed $20\ \mu\text{m}$. The γ -solid solution areas contain small fractions of segregations, likely consisting of topologically close-packed (TCP) phases.

The morphological alterations within the γ' -phase induce a transformation from lamellar and sharply-angular particle clusters (typically found in alloys produced without electrode oscillations, Figs. 3, *a* and 4, *a*) into layered texture regions when oscil-

lations occur at $f = 1.3\ \text{Hz}$. Each layer within these regions exhibits a distinct crystallographic orientation. The weld bead created with an oscillating electrode is notably wider, ranging from 150 % to 170 % wider (as shown in Fig. 2, *b*), yet its μ_b value ($0.72 < 0.88$) is lower in comparison to the bead produced without oscillations.

Increasing frequency f to 2 Hz results in the anticipated reduction of both the weld pool and crystallization front lengths. This is corroborated by more similar heating and cooling curves observed at the thermocouple locations. The cooling rate decelerates, and the temperature peaks resulting from repeated heating due to arc oscillations become less pronounced (Fig. 2, *c*). These changes lead to the predominance of the γ -solid solution in the structure, with these regions interspersed by areas containing fragmented eutectic nickel-aluminum (Fig. 3, *c* and Fig. 4, *c*). The nickel-to-aluminum ratio in these regions falls below the stoichiometric ratio required for the formation of the γ' - Ni_3Al -phase (Fig. 4, *c*). The weld bead formed at $f = 2\ \text{Hz}$ exhibits a slightly higher weld bead shape factor (0.75) while maintaining a comparable width.

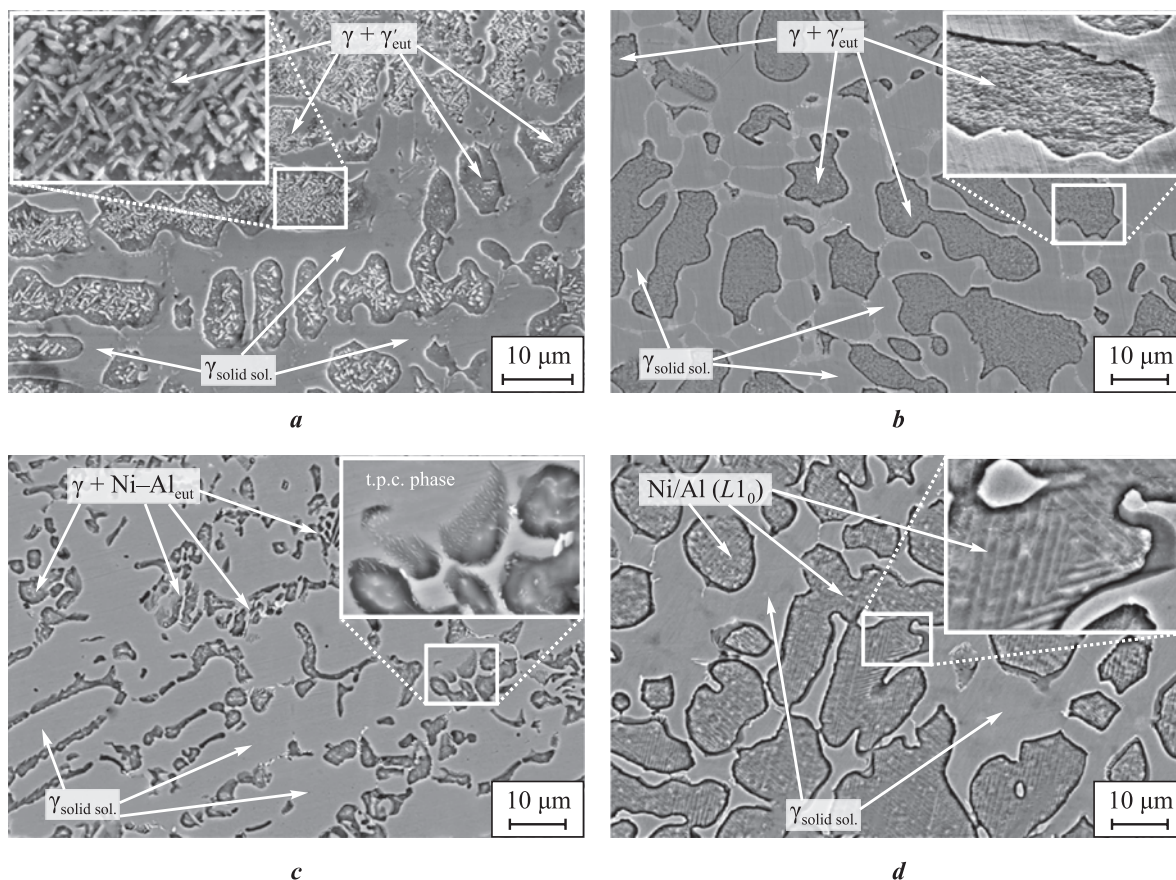


Fig. 3. Structures of the alloys surfaced without (*a*) and with electrode oscillations (*b–d*) at $f = 1.3\ \text{Hz}$ (*b*), $2\ \text{Hz}$ (*c*), and $2.8\ \text{Hz}$ (*d*)

Рис. 3. Структуры сплавов, наплавленных без колебаний электрода (*a*) и с перемещениями дуги по поверхности сварочной ванны (*b–d*) с частотой $f = 1,3\ \text{Гц}$ (*b*), $2\ \text{Гц}$ (*c*), $2,8\ \text{Гц}$ (*d*)

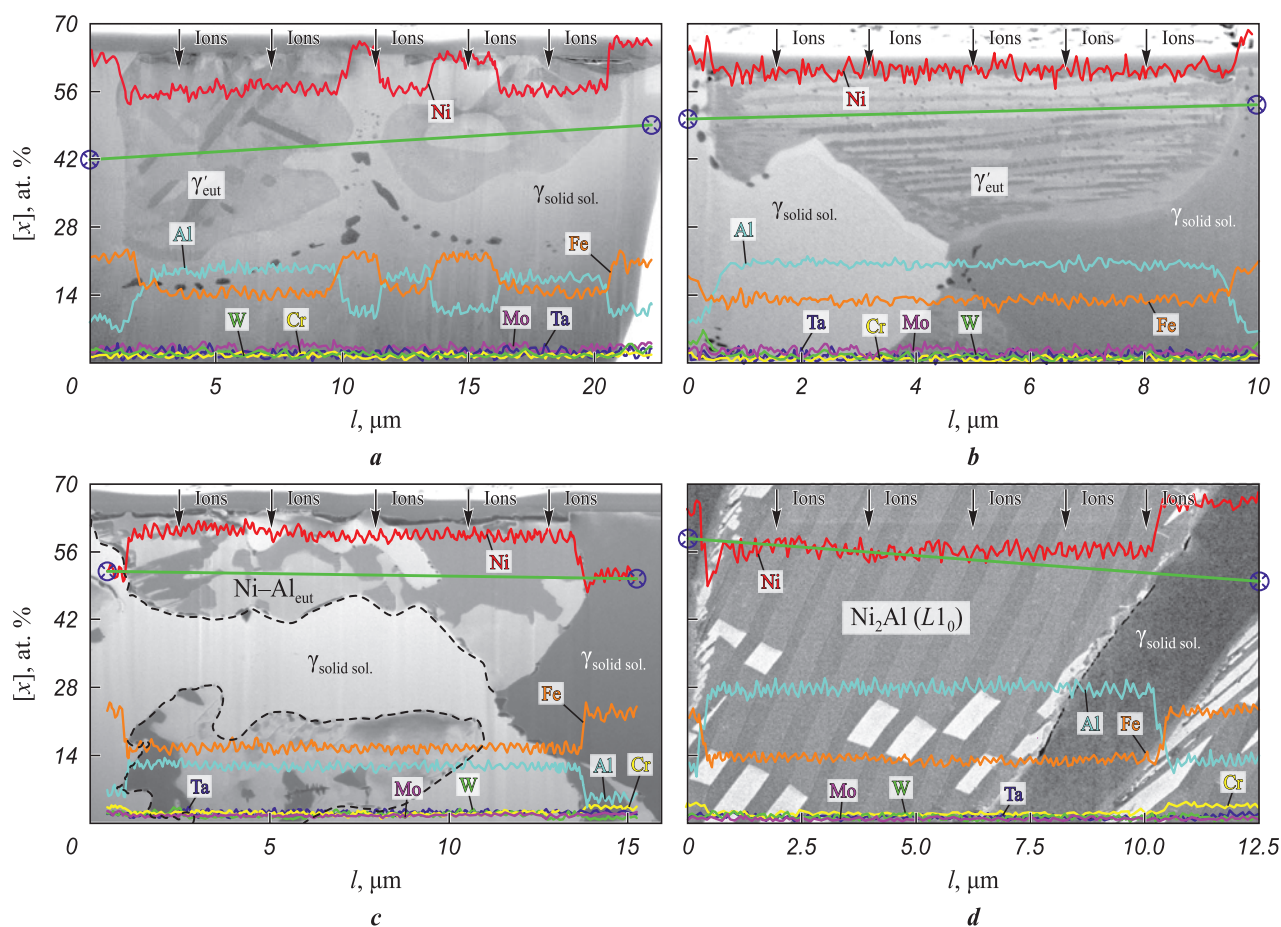


Fig. 4. Structure and elemental composition of the eutectic concentration areas after ion “etching” in the alloys surfaced without (a) and with electrode oscillations (b–d) at $f = 1.3$ Hz (b), 2 Hz (c), and 2.8 Hz (d)

Рис. 4. Структура и элементный анализ участков сосредоточения эвтектики после ионного «травления» в сплавах, наплавленных без колебаний электрода (a) и с перемещениями дуги по поверхности сварочной ванны (b–d) с частотой $f = 1,3$ Гц (b), 2 Гц (c), 2,8 Гц (d)

Further increase of the arc oscillation frequency to $f = 2.8$ Hz results in the smoothing of the temperature peaks caused by reheating in the cooling curves (as observed in Fig. 2, d) and shortens the time intervals between these peak temperatures. This phenomenon suggests a more uniform temperature gradient across the width of the weld pool. Simultaneously, the weld pool length decreases, while its volume increases due to the rapid transverse oscillations of the arc at approximately ~ 44 mm/s. Consequently, the weld bead becomes even wider ($\mu_p = 0.82$) (Fig. 2, d), in comparison to beads produced at at lower f .

At an oscillation frequency of $f = 2.8$ Hz, the conditions governing heat input to the crystallization front change, resulting in an extended period during which the crystallizing metal is at the nickel aluminide melting point (Fig. 3, d). In such conditions, the formation of the structure initiates with the emergence of alloyed γ -solid solution dendrites and concludes with the appearance of nickel-aluminum martensite ($L1_0$ lat-

tice) in the phase diagram region situated between the γ' -phase and the β -(NiAl)-phase. The primary reason for this martensite formation is the relatively slow cooling rate of the nickel-aluminum alloy from temperatures below 1200°C , in contrast to the cooling rate typical of lower frequencies in non-oscillatory surfacing. A similar hypothesis was put forth by Kositsyn S. et al. [25]. Prolonged exposure of the metal to elevated temperatures results in annealing. Prior to the formation of cooling martensite, the alloy's composition undergoes modification: it becomes enriched with aluminum atoms, which, as indicated by Kablov D. et al. [26], exhibit the highest diffusion coefficient in nickel (in the context of the alloying system under consideration) at temperatures exceeding 900°C . This shift in composition moves the alloy towards the region where the β -phase is stable. Examination of the martensitic region following surface ion etching reveals the presence of relatively thin plates (measuring 300–600 nm) with twinning orientation relative to each other (as

depicted in Fig. 4, d). The atomic mass ratio of nickel to aluminum corresponds to that of the Ni_2Al phase.

The alteration in the metal's structure due to transverse weld arc oscillations has an impact on its hardness, which varies within the range of 27 to 35 HRC depending on the oscillation frequency. It's worth noting that after ~50 thermal cycles (Fig. 5), the hardness of the alloy deposited at $f=2.8$ Hz no longer depends on subsequent thermal cycles. It remains constant, which can be attributed to the overall high thermal stability of the alloy structure.

The alloy surfaced at $f=2.8$ Hz exhibits the highest hardness value (~35 HRC), whereas the lowest hardness (27 HRC) is observed in the surfacing conducted without transverse arc oscillations (see Fig. 5). Under identical test conditions, the hardness of the commercially available Ni alloy Hastelloy C steadily decreases due to its progressive softening and loss of thermal stability.

The cobalt-based alloy Stoodite 6, characterized by a slightly higher initial hardness of 42 HRC, demonstrates robust resistance to softening. It retains its hardness almost consistently for up to 100 thermal cycles. With a further increase in the number of thermal cycles, the hardness gradually decreases and approaches the level of approximately 35 HRC, which is similar to the hardness of the alloy manufactured at $f=2.8$ Hz.

The results of thermal fatigue tests indicate that the alloy produced at 2.8 Hz exhibits the highest durability, enduring for up to 200 cycles (Fig. 5). This durability can be attributed to the attainment of an optimal ratio between the γ -solid solution and nickel-aluminum cooling martensite within the structure of a

relatively ductile metal. This metal is highly alloyed, containing up to 18 wt. % of iron and other elements. Au Y. et al. [27] reported that such an alloy is susceptible to thermoelastic transformations. During the tests, reheating of the martensite restored the reversibility of the martensite transformation. However, subsequent multiple thermal cycles lead to the dispersive decomposition of the metastable Ni_2Al -phase into Ni_5Al_3 -phase particles and a decrease in the critical temperatures of the martensitic transformation. The tungsten and tantalum contents in the examined alloy do not significantly differ, whereas chromium and molybdenum are primarily dissolved in the γ -solid solution.

The assessment of oxidative wear resulting from thermal fatigue tests indicates that the studied alloy, manufactured at $f=2.8$ Hz, exhibits superior properties compared to other available heat-resistant materials after 125 thermal cycles (Fig. 6). This difference is particularly pronounced when compared to Hastelloy C, which displays lower oxidation resistance. The diminished oxidation resistance of Hastelloy C can be attributed to the increased diffusion of oxygen from the oxidizing atmosphere through the Cr_2O_3 oxide layer. In contrast, the elevated chromium content in the alloy forms a barrier that initially hinders oxygen diffusion during the test. This barrier thickens over the course of 100–110 thermal cycles (Fig. 6). Subsequently, the protective layer experiences partial degradation.

The mass loss due to oxidation (Δm) consistently decreases in the alloy surfaced without arc oscillations. However, in the case of the alloy surfaced with arc oscillations, the mass loss remains relatively stable after 120 thermal cycles. This stability is attributed to the formation of an oxide layer on the substrate, cons-

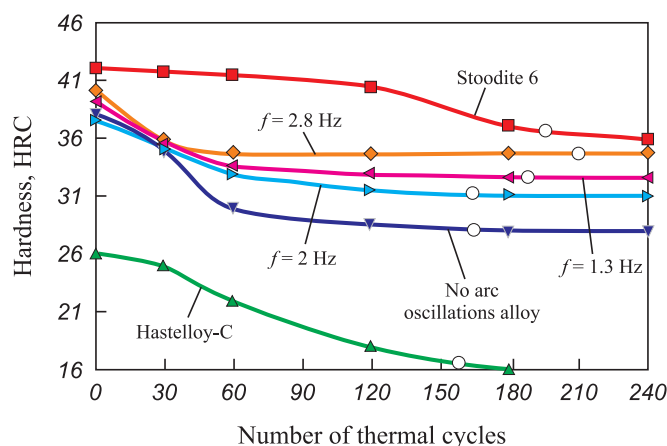


Fig. 5. Alloy hardness vs. number of thermal cycles

○ – indicates the formation of the first thermal fatigue cracks

Рис. 5. Зависимость твердости исследуемых сплавов от количества теплосмен

○ – обозначение момента, соответствующего образованию первых трещин термической усталости

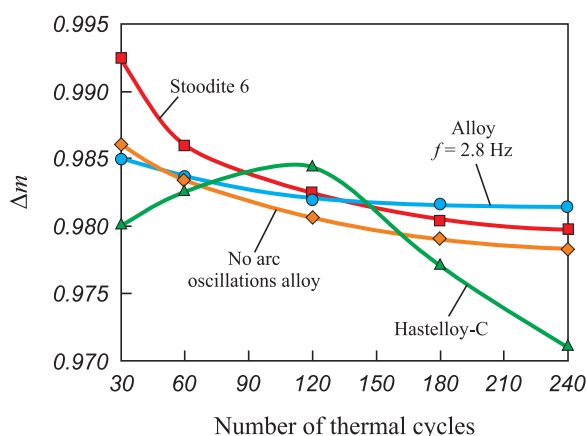


Fig. 6. Weld metal oxidation resistance vs. the number of thermal cycles

Рис. 6. Зависимость стойкости наплавленного металла к окислительному износу от количества теплосмен

tituting not less than 70 vol. % Al_2O_3 , along with traces of Cr_2O_3 and 5 vol. % NiO. The cobalt alloy Stoodite 6 demonstrates reasonably high oxidation resistance due to the formation of the CoCr_2O_4 oxide. It only becomes less effective than the studied alloy after 125 thermal cycles when oxidation microcracks intensively develop in the regions susceptible to oxidation of Cr_7C_3 [28]. Within the margin of measurement error, it can be inferred that the oxidation resistance of the studied alloy is comparable to that of Stoodite 6.

The surfacing process involving weld arc oscillation effectively regulates the heat input to the crystallization front, thereby establishing favorable conditions for the development of coatings with exceptional resistance to thermal fatigue cracks.

Conclusion

1. Transverse arc oscillations applied to the weld pool surface change the configuration of the crystallization front and the thermal conditions governing the formation of the alloyed nickel aluminide structure. At a low electrode oscillation frequency ($f = 1.3$ Hz), a structure is generated in which the clusters of γ' - Ni_3Al phase lamellar and sharp-angular particles transform into layered regions. This transformation enhances the resistance to thermal fatigue cracking compared to the alloy surfaced without electrode oscillations.

2. In order to achieve the utmost resistance to thermal fatigue cracking, the thermal conditions (at 2.8 Hz) should promote gradual cooling of the nickel-aluminum alloy and the development of a balanced structure comprising the γ' -solid solution, heavily alloyed (up to 18 wt. %) with iron and other elements, as well as nickel-aluminum martensite. The composition of this martensite corresponds to the metastable Ni_2Al phase, and reheating it does not render the alloy more brittle.

3. The weld pool temperature gradient stabilizes at $f = 2.8$ Hz, concurrently leading to a reduction in the length of the weld pool and the attainment of the highest weld bead shape factor ($\mu_b = 0.83$). Consequently, the extent of bead-to-bead overlap can be decreased to 20–30 %, thereby reducing the consumption of relatively expensive surfacing material.

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<https://doi.org/10.1016/j.msea.2017.12.057>

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С. А. Фастов – участие в изготовлении композиционной проволоки для наплавки и в экспериментах по ее наплавке.

Received 17.05.2023

Revised 14.07.2023

Accepted 18.07.2023

Статья поступила 17.05.2023 г.

Доработана 14.07.2023 г.

Принята к публикации 18.07.2023 г.